Monday, March 2nd:

A Moment of Science

* If you didn’t turn in the frog sheet, it’s due now (and it’s late [except 7th period])
* If you were absent Friday AMOS 2/17 to 2/28 needs to be turned in now.

**Reminder:** you have a vocab quiz 3/6! Start studying now! Multiple choice 16 questions (30 words of 40)

Write “quiz” on your paper.

Take a notecard from Mrs. Lewis, and answer the following:

1. What are the 7 diatomic elements?

*Bonus (3 points):* How do you write a diatomic element?

Then use your textbook or notes to draw a depiction of equally shared electrons and unequally shared electrons (dipole). Use specific examples (hint, think about ∆E).

**Please note, you are responsible for ensuring your surrounding area is left neat and tidy. If you do not, you will be asked to write the following at least 10 times.**

I am a capable young adult, and Mrs. Lewis is NOT my maid. I will ensure my table and surrounding area is clean before I leave the room daily.

Tuesday, March 3rd:

A Moment of Science

1. In your own words describe the process of writing Lewis dot structures.
2. In general, how many electrons does each atom want around it in a Lewis dot structure?
3. Number 2 above is an example of the \_\_\_\_\_\_\_\_\_ rule.
4. Using Lewis dot structures, show how lithium donates an electron to chlorine in an ionic reaction.
5. Using Lewis dot structures, show how carbon bonds with oxygen covalently to make carbon dioxide, CO2.

Wednesday, March 4th/Thursday, March 5th:

A Moment of Science

1. Covalent compounds are made up of what types of elements? Two nonmetals
2. What are the prefixes for covalent compound naming? Mono di tri tetra penta hexa hepta octa nona deca
3. Do you use mono- on the first element listed in a covalent compound? No, only on the 2nd element
4. When naming, how do you change the ending of the 2nd element in the covalent molecule? Add -ide ending
5. Name the following covalently bonded molecules.
6. N2O4  dinitrogen tetroxide
7. CO2 carbon dioxide
8. CO carbon monoxide
9. N2O dinitrogen monoxide
10. CCl4 carbon tetrachloride
11. Draw the Lewis dot structures of the molecules in number 4 b, c, d, e above. Circle the electrons that get shared when bonding.

Friday, March 6th:

A Moment of Science

Write “Vocab quiz” on your paper.

Take 5 min to study your vocab words. Words on pg 195 are NOT on the quiz. It is MATCHING and 16 questions.

When you finish, put into the tray and prepare to take notes on your AMOS sheet on the video below.

<https://www.youtube.com/watch?v=cIuXl7o6mAw>

Monday, March 9th:

A Moment of Science

* If you were absent Wed/Thurs/Fri, know that you have make-up work including a vocab quiz. Arrange a time with Mrs. Lewis to complete such. (your study hall, 3rd, tutoring Tuesday 3:30)
* You can build the molecules/compounds with fruit loops from the lab if you want, but the **lab paper MUST BE done** as it was for a grade.
* Bring your books on Tuesday! Book check for a grade!

As there was a lot of people out, we will re-watch this video <https://www.youtube.com/watch?v=cIuXl7o6mAw>

If you were out, please take notes! If you already watched it and have notes on it, please sit quietly and reabsorb the material!

Tuesday, March 10th:

A Moment of Science

Textbook checks! Put them on my table and I’ll grade you for having them.

1. Ionic naming
	1. If made up of 2 elements, what do you change the 2nd element ending to? -Ide
	2. If includes a 3+ elements you have a polyatomic ion in your compound. What endings are typical of polyatomics? -ate and -ite
	3. Which comes first, cation or anion? Cation
	4. When do you need parenthesis? To show more than one polyatomic element; when you move a number (other than 1) behind a polyatomic
	5. What type of metals would include Roman numerals in their name? transition metals because they have variable charges
	6. Why is it important to include Roman numerals? It tells you the charge of your transition metal ion

Wednesday, March 11th/Thursday, March 12th:

A Moment of Science

1. What page in your textbook are the polyatomics listed? Page 210

Turn to that page and answer the following questions. You may also want to use page 205 to assist.

1. What is the formula of Copper(I) sulfate? Will you need parenthesis? Why/Why not? Cu2SO4 no because copper’s charge of 1 goes behind the polyatomic, so there’s no need for parenthesis
2. What is the formula of Magnesium bicarbonate? Will you need parenthesis? Why/Why not? Mg(HCO3)2 yes, because magnesium’s charge of 2 goes behind the polyatomic and you’ll need to separate the 3 from the 2
3. What is the name of Pb(NO3)4 Lead(IV) Nitrate
4. What is the name of Na3N Sodium Nitride
5. All of the above are what types of bonds? Ionic (metal + nonmetal)

Friday, March 13th:

A Moment of Science

In your own words, describe the process of determining the name of an ionic compound. Use MgCl2 Ca(CO3)2 and NaNO3 as examples.

In your own words, describe the process of determining the formula of an ionic compound. Use magnesium phosphate, lead(II) chloride, and beryllium carbonate as examples.